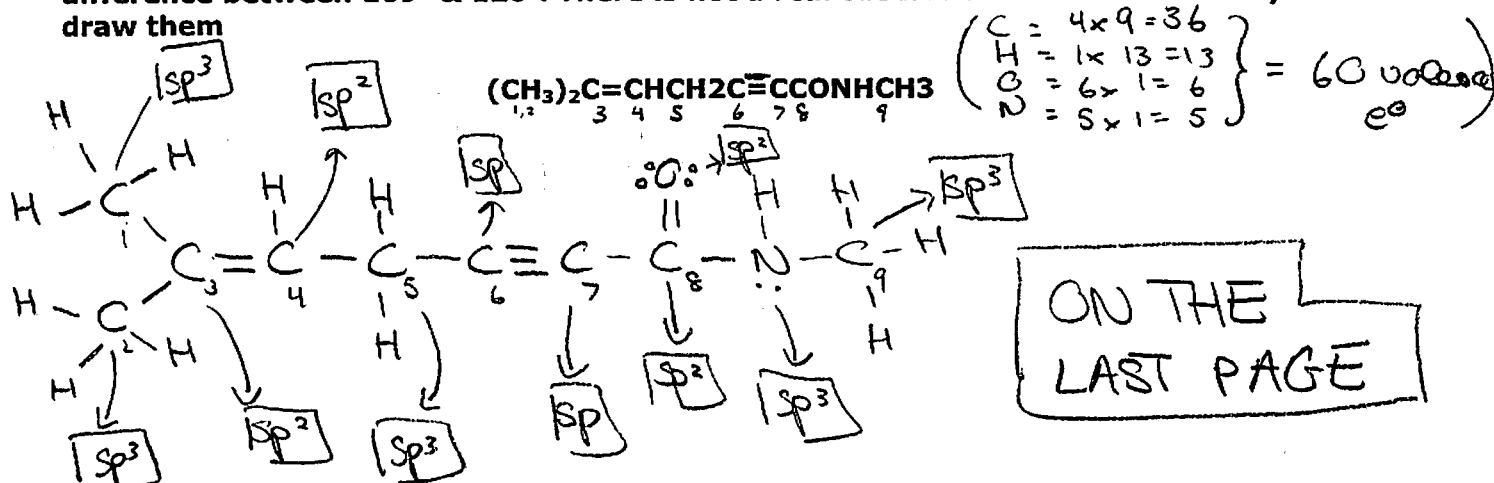


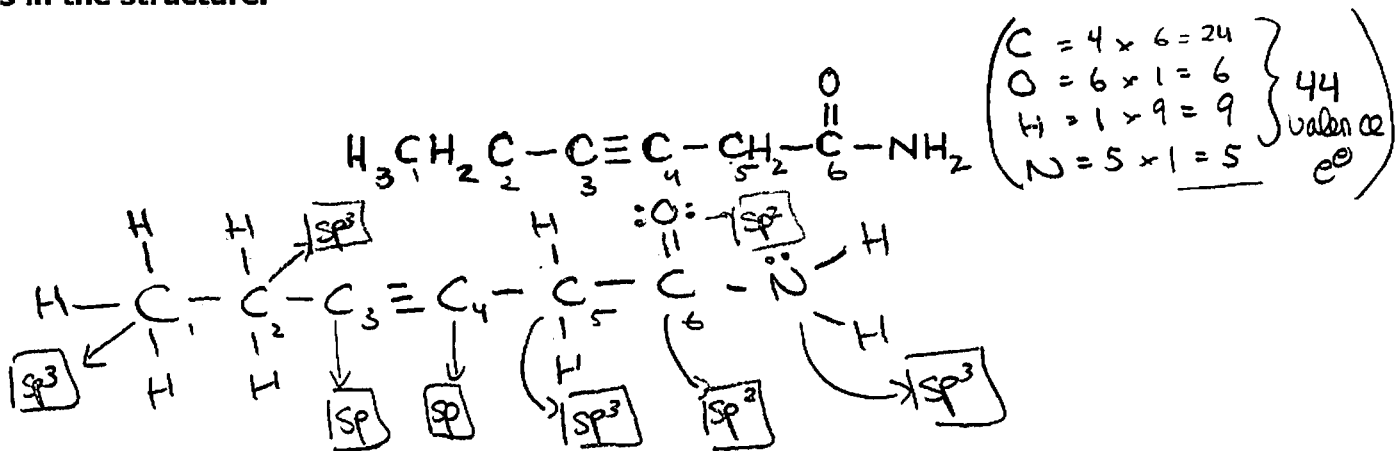
Name: KEY Date: _____ Section: _____
 Applying Geometry to Hybrid Orbitals- graded classwork

Rules:
 sp³ hybridized C- 109° - tetrahedral
 sp² hybridized C- 120°- trigonal planar
 sp hybridized C- will involve pi bond- linear

Draw an accurate geometric representation of the following structure. Keep the carbon backbone in same plane, including the N. The other atoms should come off in front of the paper and behind the paper (dotted line). Don't worry to much about the difference between 109° & 120°. There is not a real observable difference when you draw them



What is the hybridization of each atom in the following molecule? You can ignore the H's in the structure.



3. Show the cloud structure of the following molecule. Label the hybridization of each carbon and show the cloud structure of molecule.

